Solvothermal synthesis of discrete cages and extended networks comprising \( \{\text{Cr(III)}_3\text{O(O}_2\text{CR})_3(\text{oxime})_3\}^{2-} \) (\( R = \text{H, CH}_3, \text{C(CH}_3)_3, \text{C}_{14}\text{H}_9 \)) building blocks†

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The synthesis, structural and magnetic characterisation of a family of related Cr(III) cages are reported. Each member comprises \( \{\text{Cr(III)}_3\text{O(O}_2\text{CR})_3(\text{R}_2\text{-sao})_3\}^{2-} \) (\( R_1 = \text{H, CH}_3, \text{C(CH}_3)_3, \text{C}_{14}\text{H}_9; R_2 = \text{Me, Ph, }^{3}\text{Bu, C}_{10}\text{H}_8 \)) triangles linked by Na\(^+\) cations, resulting in either the discrete complexes \( \{\text{H}_3\text{O}\}_2\{\text{NaCr(III)}_6\text{O}_2(\text{O}_2\text{CCH}_3)_6(\text{Me-sao})_6\} \) (1) and \( \{\text{Na}_4\text{Cr(III)}_6\text{O}_2(\text{O}_2\text{CCH}_3)_6(\text{3,5-di-}^{3}\text{Bu-sao})_6(\text{MeCN})_6\} \) (3); or the extended networks \( \{\text{H}_3\text{O}\}_2\{\text{Na}_2\text{Cr(III)}_6\text{O}_2(\text{Me-sao})_6(\text{Ph-sao})_6(\text{MeCN})_2(\text{H}_2\text{O})_2\}_n \) (4) and \( \{\text{Na}_2\text{Cr(III)}_6\text{O}_2(\text{CH}_3\text{C(OH)}_3)_6(\text{Me-sao})_6(\text{MeCN})_2(\text{H}_2\text{O})_2\}_n \). Magnetic susceptibility data obtained for 2 and 4 reveal weak antiferromagnetic exchange between the Cr(III) ions in the triangles.

Results and discussion

Pertinent crystallographic data for all five complexes are given in Tables S1 and S2.† All metal oxidation states were

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† Electronic supplementary information (ESI) available. CCDC 1481774–1481778. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6ra14811e
confirmed using bond valence sum (BVS) calculations and bond length and charge balancing considerations. The aerobic stirring of an acetonitrile solution of CrCl₃·6H₂O with equimolar amounts of Naphth-saoH₂, sodium-9-anthracene carboxylate and base [NEt₄][OH] affords a green slurry. When this slurry is heated solvothermally for a period of 24 hours at T = 100 °C, followed by cooling at a rate of 3 °C per hour, crystals of the heptanuclear complex [H₂O][NEt₄]₂·[Na₁Cr(m)₆(O)₂(O₂C-C₁₄H₉)₆(Naphth-sao)₆] (1) are obtained. Longer cool down periods (range tested: 1–3 °C per hour) have no significant effect on yield and crystal quality. Complex 1 crystallises in the monoclinic space group C2/c, with one half [Na₁Cr(m)₆] unit in the asymmetric unit. The core in 1 has a sandwich-like shape where a central Na⁺ ion (Na1) lies at the midpoint of two \{Cr(m)₃O(O₂C-C₁₄H₉)₆(Naphth-sao)₆\}²⁻ moieties in 1 (and in 2–5) have been previously observed only in Mn(m) coordination chemistry,¹⁵ and are adaptations on the classic and ubiquitous trinuclear basic carboxylates [Cr₃O(O₂C-CN)₆(L)₆]²⁻ (L = H₂O, pyridine etc.),¹⁶ whereby the \{Cr(m)₃O\}²⁺ core is maintained while substituting half the carboxylates for three Naphth-sao²⁻ ligands to give the structure in 1 (Fig. 1). Each edge of these trinuclear fragments are bridged by one μ-O-C-anthracenoate (O₂C-C₁₄H₉) and one η¹:η²:η³, μ₁-Naphth-sao²⁻ ligand which occupy opposite sides of the \{Cr(m)₃O\}²⁺ plane (Fig. 1 and 2). The O donor atoms of the carboxylates exclusively bridge the Cr(m) centres within the triangular units. The Naphth-sao²⁻ ligands also bridge between neighbouring Cr ions, but also connect the \{Cr(m)₃O\}²⁺ units to the central Na⁺ ion via six oximic O-atoms (O₁₂, O₁₄, O₁₆ and symmetry equivalent, s.e.) with distances lying in the 2.377(5)–2.442(5) Å range (Fig. 2). The resultant \{Na₁Cr(m)₆(O)₂(O₂C-C₁₄H₉)₆(Naphth-sao)₆\}³⁻ anionic clusters are charge balanced by a hydronium ion and two \{NEt₄\}⁺ counter ions.

The \{NEt₄\}⁺ counter anions in 1 sit within pockets formed by two neighbouring \{NaCr₆\} units, and are held in position by numerous C–H⋯π interactions between ethyl protons and the surrounding aromatic rings of the Naphth-sao²⁻ and O₂C-C₁₄H₉ ligands (e.g. C₈S₅(H₈SB)⋯[C₁₅-C₂₀]centroid = 2.646 Å and C₈S₃(H₈SB)⋯[C₆₆–C₇₇]centroid = 2.803 Å; Fig. S1†). The \{NaCr₆\} units arrange into 2D brickwork sheets (bc plane) which are stacked in parallel rows along the a unit cell axis (Fig. S2†).

The solvothermal reaction of CrCl₃·6H₂O, Ph-saoH₂, sodium formate and tetraethylammonium hydroxide in MeCN, affords the extended network [H₂O]₂[Na₂Cr(m)₆O₂(O₂CH)₆(Ph-sao)₆(MeCN)₃(H₂O)]·4MeCN (2) (Fig. 3). Complex 2 crystallises in the monoclinic space group P2₁/n and its core comprises two triangular \{Cr(m)₃O(O₂CH)₆(Ph-sao)₆\}²⁻ units that, unlike in 1, are linked via two central distorted octahedral Na⁺ ions (Na1 and s.e.). An inversion centre lies at the midway point between the alkali metals. The coordination spheres at the two central Na⁺ centres are each completed by terminal H₂O and MeCN ligands (Na₁–N₄ = 2.441(12) Å, Na₁–O₁₄ = 2.328(7) Å). The six carboxylates in 2 employ a combination of μ- and η¹:η², μ₁-bonding motifs to fuse the central Na⁺ ions (Na1 and s.e.) to the \{Cr(m)₃O\}²⁺ units, while also linking the Cr(m) centres within the triangular moieties. Similarly, the six Ph-sao²⁻ ligands exhibit both η¹:η¹:η¹, μ₂ and η⁵:η¹:η¹, μ₄-bonding modes to connect the Na⁺ ions to the \{Cr(m)₃\} triangular units (Fig. 3b). Two symmetry equivalent hydronium ions (O15 and s.e.) lie juxtaposed to the core in 2 and are held in position through H-bonding interactions with O-donor atoms (O₃, O₉–O₁₁) of nearby formate and Ph-sao²⁻ ligands (O₁₅⋯O₃ = 2.718 Å, O₁₅⋯O₉ = 2.921 Å, O₁₅⋯O₁₀ = 2.835 Å and O₁₅⋯O₁₁ = 2.919 Å) and two MeCN solvent of crystallisation (N₅ and N₆) (O₁₅⋯
N5 = 3.077 Å, O15···N6 = 2.935 Å. In combination with the linker Na⁺ ions, these hydronium ions effectively connect the individual [Cr₃] units in 2 towards the formation of wave-like 2D sheets that intersect the bc plane of the unit cell and stack upon one another along the a cell direction in a space efficient parallel fashion (Fig. 4b and S3†). The overall result is an [3,6] network topology in 2 as shown in Fig. 4a.

Employment of the ligand 3,5-di-tert-butyl-salicyaldoxime in combination with pivalate co-ligands in an analogous reaction to those described above gives rise to the complex [Na₄Cr(III)₆(O)₂⁻(O₂CC(CH₃)₃)₆(3,5-di-tBu-sao)₆(MeCN)₆] (3). Complex 3 crystallises in the trigonal space group R̅3c and bears a similar sandwich-like core to that observed in 1 (compare Fig. 5c with 2). However in this particular case the two {Cr(III)₃O(O₂CC(CH₃)₃)₃(3,5-di-tBu-sao)₃}{O₂CC(CH₃)₃}₂⁻ moieties are connected through a belt of four Na⁺ ions (Na1–2 and s.e.). Na2 (as in 1) is bound by oximic O-atoms (O2, O6 and s.e.) and lies at the centre of the cage, equidistant between the two central O₂⁻ ions (O5 and O7) located at the centre of each {Cr(III)₃O}⁷⁺ unit. The three symmetry related sodium centres (Na1 and s.e.) are situated in the same plane as the central Na2 and together form a star shaped arrangement (Fig. 5a). The coordination spheres at Na1 (and s.e.) are completed by two terminal MeCN ligands (Na1–N2 = 2.497(7) Å, Na1–N3 = 2.454(7) Å) (Fig. 5b and d). As observed in 1, the carboxylates in 3 bridge only the Cr(III) ions within each trinuclear unit using the common μ-bridging mode. As a result the ‘Bu-sao²⁻’ ligands must assume the role of connecting the {Cr(III)₃O}⁷⁺ units to the central belt of Na⁺.
ions and this is achieved using a $\eta^2: \eta^1: \eta^3$, $\mu_5$-bonding motif (Fig. S4†). On moving from complex 1 to 3, it becomes apparent that the number of Na" ions increases from one (in 1) to four (in 3); this may tentatively be assigned to the presence of less sterically demanding organic ligands. In the crystal of this may tentatively be assigned to the presence of less sterically

A combination of the Me-saoH$_2$, acetate ligands and NaOH leads to the formation of [H$_2$O][Na$_3$Cr(III)$_6$(O)$_2$(O$_2$CCH$_3$)$_6$(Me-sao)$_6$]$_n$[MeCN]$_m$ (4) (Fig. 7). Complex 4 crystallises in the monoclinic $P2_1/a$ space group and is the second example of an extended network in this family of compounds. Its structure comprises [Cr(III)$_3$(O)(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ nodes linked into a 2D [3,6], sheet-like topology via charge balancing Na" ions. The connectivity within each [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ unit is as follows: the three Me-sao$^{2-}$ ligands located within each [Cr$_3$] moiety exhibit the $\eta^5: \eta^1: \mu_5$-bonding mode and each span an edge of a trinuclear unit while also providing a connection to the Na" ion nodes (Na1–2 and s.e.). An identical role is adopted by the three acetate ligands within each [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ unit in 4, using a combination of $\eta^2: \eta^1$, $\mu_5$- and $\eta^5: \mu_4$-bonding motifs (Fig. S5†). As a result the central [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ nodes are connected into 2D sheets via three bridging Na" ions, resulting in a [3,6] topology as shown in Fig. 7b. The hydronium ions (O14 and s.e.) sit directly above the [Cr$_3$] units in 4 and stitch together the individual 2D sheets through H-bonding interactions (e.g. O14…O5 = 2.873 Å and O14…O7 = 2.901 Å). The extended connectivity in 4 is also enhanced by numerous inter-sheet C–H…π interactions between aromatic oxime and neighbouring acetate protons (C$_2$(H$_2$B)…[C$_9$–C$_{14}$]centroid = 3.381 Å and C$_2$(H$_2$O)…[C$_1$–C$_r$]centroid = 2.796 Å). Na1 exhibits a distorted octahedral geometry while Na2 possesses a distorted square-based pyramidal configuration ($\tau = 0.10$). A single acetonitrile ligand completes the coordination sphere at Na1 (Na1–N4 = 2.437(7) Å).

A repeat of the reaction that produced 4, but under ambient conditions affords green crystals of the species [Na$_3$Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$[H$_2$O]$_{3\cdot}$3MeCN (5) after the methanolic slurry is taken to dryness and recrystallised from MeCN (trigonal R$3$ space group). Akin to 4, complex 5 also exhibits a 2D [3,6] extended network topology and a core comprising [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ nodes linked by Na" ions (Na1 and s.e.) to give a highly symmetrical hexagonal honeycomb array (Fig. 8 and 9). To create this 2D topology the Me-sao$^{2-}$ and O$_2$CCH$_3$ ligands each bridge four metal centres ($2 \times$ Cr(III) and $2 \times$ Na") by adopting $\eta^5: \eta^1: \mu_5$ and $\eta^5: \mu_4$-bonding modes, respectively (Fig. S6†). Two terminal H$_2$O ligands (O6 and O7) bond to each sodium linker ion (Na1–O6 = 2.372(3) Å, Na1–O7 = 2.373(4) Å) and in doing so lie within the hexagonal cavity while effectively completing their distorted octahedral geometries. The individual 2D sheets in 5 pack in

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Fig. 7 (a) A single [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ node surrounded by three connecting Na" ions as viewed perpendicular to the [Cr(III)$_3$(O)$_2$]$_n^+$ plane in 4. (b) The 2D [3,6] network in 4 where the red vertices represent the O (O1) atom at the centre of the [Cr(III)$_3$(O)$_2$]$_n^+$ triangular nodes and the yellow vertices represent the bridging Na" ions (Na1–2). (c) Colour coded and space-fill represented 2D nets in the unit cell of 4.

Fig. 8 A single [Cr(III)$_3$(O)$_2$]$_n^+$ unit along with the linker Na" ions (yellow spheres) as viewed perpendicular (a) and parallel (b) to the triangular plane in 5. (c) An array of six [Cr(III)$_3$(O)$_2$(O$_2$CCH$_3$)$_3$(Me-sao)$_3$]$^{2-}$ units linked by Na" ions to form the hexagonal [3,6] honeycomb topology in 5.
a parallel fashion along the c direction of the unit cell as illustrated in Fig. 10 (inter-sheet distance: Na1⋯Na1′ = 12.272 Å). Moreover these individual sheets arrange in a staggered parallel arrangement in relation to one another and therefore no significant pores or channels are observed in 5. However, MeCN solvent of crystallisation occupy the spaces in-between each of the 2D sheets and do so by arranging themselves into symmetry related triads around the three fold axis of the cell (down c), while also H-bonding with terminal H2O ligands (N2⋯O6 = 2.845 Å) (Fig. 9).

TGA-DSC measurements

The TG trace of [Na4Cr(III)6O2(O2CC(CH3)3)6(3,5-di-t-Bu-sao)6(MeCN)6](3) shows two distinct mass loss regions, with the initial loss of 10.4% corresponding to all (7.5 per cage) coordinated and interstitial MeCN units (calculated as 11.1%). The second (and much steeper) mass drop occurs at ~420 °C and is attributed to the sublimation of 3 (or its related decomposition products). The corresponding trace for [Na2Cr3O(O2CC(CH3)3(Me-sao)3(H2O)6]$\cdot$3MeCN (5) exhibits three mass loss regions. The initial loss (in the 50–100 °C range) of 18.1% is attributed to the removal of interstitial MeCN solvent molecules and terminal H2O solvent ligands (calculated at 21.3%). The closely related second and third steps between 300 and 475 °C represent an 18% loss and is tentatively attributed to acetate ligand loss (calculated as 16.3%), which is rapidly followed by further ligand loss and eventual decomposition of 5 (Fig. S7†).

Magnetic susceptibility measurements

The d.c. molar magnetic susceptibility, $\chi$, of polycrystalline samples of 2 and 4 were measured in an applied magnetic field, $B$, of 0.1 T, over the 5–300 K temperature, $T$, range (Fig. 11, where $\chi = M/B$, and $M$ is the magnetisation). Both show very similar behaviour, and we discuss the data per {Cr(III)3} triangle assuming no interaction through the Na+ ions. At room temperature, the $\chi T$ products of both are approximately 5.5 cm$^3$/mol·K, close to the value expected from the spin-only contribution to the magnetism of a trinuclear Cr(III) unit (5.6 cm$^3$/mol·K, with $g_{Cr} = 2.00$). Upon cooling, the $\chi T$ products decrease continuously to reach a value of approximately 1.8 cm$^3$/mol·K per [Cr(III)3] at 5 K. This behaviour is indicative of weak antiferromagnetic interactions between the Cr(III) ions. For the interpretation of the magnetic properties of 2 and 4, we consider that they arise from an equilateral triangle with just one exchange interaction (Fig. 11-inset). Thus we used spin-Hamiltonian (1):

$$\hat{H} = \sum_{i<j} -2J_{ij}\hat{S}_i\hat{S}_j + \mu_B B_{GC\text{r}} \sum_i \hat{S}_i$$  (1)
Concluding remarks

The solvothermal heating of Cr(III) salts in the presence of bulky R-saoH₂ oxime ligands (Naphth-saoH₂ and Ph-saoH₂) with carboxylate co-ligands of varying sizes produces a series of structurally related discrete cages each comprising Na⁺ linked [Cr(III)₃O(O₂CR)₃(R-sao)]³⁻ units. Siblings 1–5 are thus new members of the very small family of Cr(III) cages stabilised with phenolic oximes, and their formation suggests many more family members await discovery through judicious ligand design and the exploration of a variety of reaction conditions. Moreover, compounds 2, 4 and 5 are the first examples of extended network materials comprising Cr-oxime building blocks. Magnetic susceptibility measurements on 2 and 4 reveal weak antiferromagnetic exchange between neighbouring Cr(III) ions within the isolated triangular units.

Experimental section

Physical measurements

Infra-red spectra were recorded on a Perkin Elmer FT-IR Spectrometer equipped with a Universal ATR Sampling accessory (NUI Galway). Elemental analysis was carried out by Marion Vignoles of the School of Chemistry microanalysis service at NUI Galway. Variable-temperature, solid-state direct current (d.c.) magnetic susceptibility data down to 5 K were collected on a Quantum Design MPMS-XL SQUID magnetometer equipped with a 7 T dc magnet (University of Edinburgh). Diamagnetic corrections were applied to the observed paramagnetic susceptibilities using Pascal’s constants. All measured complexes were set in eicosane to avoid torquing of the crystallites. All magnetic samples are collected as single-crystalline products and analysed using microanalysis and IR measurements prior to their magnetic assessment. If necessary, phase purity between cross-batches are validated using unit cell checks and IR measurements. TGA measurements were carried out by Dermot McGrath (NUI Galway) using a Rheometric Scientific STA 625.

X-ray crystallography

The structures of 1–5 were collected on an Xcalibur S single crystal diffractometer (Oxford Diffraction) using an enhanced Mo source. Each data reduction was carried out on the CrysalisPro software package. The crystal structures of 1 and 3 was solved by an dual-space algorithm using SHELXT,²⁰ whereas 2, 4 and 5 were solved by direct methods using SHELXS-97 (ref. 21) and refined by full matrix least squares using SHELXL-2014 (ref. 19) within OLEX2 (ref. 21) (structures 1 and 3) and SHELXL-97 (ref. 20) (structures 2, 4 and 5) within OSCAIL software package.²² All hydrogen atoms were placed in calculated positions. All non-hydrogen atoms were refined as anisotropic. Despite numerous attempts each single crystal data set obtained from 1–3 were found to consistently diffract poorly at higher angles. Our best data sets have been supplied in this work. Residual electron densities in solvent accessible voids and channels were observed in 1, 3 and 5 and so were modelled using the SQUEEZE program.²³ The two large channels (volumes ~ 970 Å³) in 1 contained extremely diffuse electron density and were assumed to contain numerous solvents of crystallisation (H₂O and MeCN) along with the required oxonium charge balancing cation. CHN analysis on 1 supported these observations. Due to poor quality of data, SIMU, DELU and RIGU restraints were applied to the crystal structure in 1. Two tert-butyl groups in 3 (at C5 and C37, respectively) were each modelled as disordered over two sites with ratios of 61.7/38.3 (at C5) and 86.5/13.5 (at C37). Bond length (C–C) and angle (C–C–C) restraints were also required to maintain sensible geometries at these sites. Likewise, global SIMU/DELU and RIGY restraints were employed to improve the atomic displacement parameters at the disordered C6 and C6A sites (see cif for full details).
Materials and syntheses

All solvothermal reactions (1–4) were carried out in a Heraeus (UT6420-Thermo Scientific) oven using spring loaded stainless steel digestion vessels (23 cm³ capacity) produced by the Parr Instrument Company. All reagents and solvents were used as purchased. The salicyaldoxime ligands were synthesised by the reaction of the precursor ketone with hydroxylamine and purchased. The salicyaldoxime ligands were synthesised by the reaction of the precursor ketone with hydroxylamine and purchased. Sodium acetate in ethanol, as described in the literature. The synthesis of chromium pivalate was adapted from established methods,23 while chromium acetate was purchased from Sigma Aldrich and used without purification. Complex 5 was carried out on the bench top under ambient conditions using solvents that were used as purchased without further purification.

[C₄H₆O][NEt₄][NaCr(m)₆(O)₃(O₂C-C₆H₄)₆(Naphth-sao)₆](Na₄Cr₆(O)₂(O₂C-C₆H₄)₆(Naphth-sao)₆) (1). CrCl₃·6H₂O (0.1 g, 0.37 mmol), Naphth-saoH₂ (0.07 g, 0.037 mmol), sodium-9-anthracene carboxylate (0.09 g, 0.037 mmol) and tetraethylammonium hydroxide (0.2 cm³, 0.15 mmol) were stirred in MeCN (10 cm³) for 6 hours. This slurry solution was then transferred to a 23 cm³ capacity Teflon vial which was then lodged in a spring loaded stainless steel digestion vessel and placed in an oven. The temperature was then lowered gradually down to room temperature over a period of 24 hours. The temperature was then lowered gradually down to room temperature over a period of 24 hours. Dark green rhombic shaped crystals of 1 were collected and air dried to give a yield of 10%. Elemental analysis (%) calculated (found) for C₁₇₃H₁₉₃N₅O₅₄Na₁Cr₆: C 51.30 (51.01), H 3.95 (3.75), N 7.48 (7.39). FT-IR (cm⁻¹): 3046(vb), 2965(w), 1589(w), 1549(s), 1481(w), 1430(vs), 1301(vs), 1244(s), 1164(w), 1020(m), 971(m), 956(w), 898(w), 872(v), 835(m), 812(w), 788(m), 746(w), 715(s), 695(m).

[N₂O₅][NaCr(m)₆(O)₃(O₂C-C₆H₄)₆(Me-sao)₆(MeCN)₂(H₂O)₂] (2). CrCl₃·6H₂O (0.1 g, 0.38 mmol), Ph-saoH₂ (0.08 g, 0.38 mmol), sodium formate (0.05 g, 0.74 mmol) and tetraethylammonium hydroxide (0.2 cm³, 0.15 mmol) were stirred in MeCN (10 cm³) for 2 hours. This slurry solution was then transferred to a 23 cm³ capacity Teflon vial which was then placed in an spring loaded stainless steel digestion vessel and placed in an oven. The oven temperature was raised to 100 °C over a period of 5 minutes and held at this temperature for 24 hours. The temperature was then lowered gradually down to room temperature over a period of 24 hours. Dark green rhombic shaped crystals of 2 were collected and air dried to give a yield of 15%. Elemental analysis (%) calculated (found) for C₃₀H₅₄O₂₅N₃Na₂Cr₃: C 41.92 (41.58), H 4.09 (4.43), N 5.52 (5.75). FT-IR (cm⁻¹): 3369(b), 2957(w), 1589(w), 1549(s), 1481(w), 1430(vs), 1301(vs), 1296(w), 1255(m), 1181(m), 1141(w), 1087(w), 1041(m), 1025(m), 951(s), 885(w), 864(w), 822(w), 785(m), 750(m), 735(s), 678(s).

[N₂O₅][NaCr(m)₆(O)₃(O₂C-C₆H₄)₆(Me-sao)₆(MeCN)₂(H₂O)₆]·4MeCN (3). CrCl₃·6H₂O (0.1 g, 0.38 mmol), Ph-saoH₂ (0.08 g, 0.38 mmol), sodium formate (0.05 g, 0.74 mmol) and tetraethylammonium hydroxide (0.2 cm³, 0.15 mmol) were stirred in MeCN (10 cm³) for 2 hours. This slurry solution was then transferred to a 23 cm³ capacity Teflon vial which was then placed in an spring loaded stainless steel digestion vessel and placed in an oven. The oven temperature was raised to 100 °C over a period of 5 minutes and held at this temperature for 24 hours. The temperature was then lowered gradually down to room temperature over a period of 24 hours. Dark green rhombic shaped crystals of 3 were collected and air dried to give a yield of 10%. Elemental analysis (%) calculated (found) for C₃₀H₅₄O₂₅N₃Na₂Cr₃: C 34.03 (33.85), H 4.09 (4.43), N 7.48 (7.76). FT-IR (cm⁻¹): 3369(b), 1587(w), 1490(w), 1471(w), 1436(m), 1394(w), 1359(m), 1311(s), 1250(w), 1155(w), 1045(m), 1026(m), 957(s), 846(s), 759(s), 673(s).

[NaCr(m)₆(O)₃(O₂C-CH₃)₆(3,5-di-‘Bu-sao)₆(MeCN)₆] (4). Chromium pivalate (0.1 g, 0.20 mmol), 3,5-di-tert-butyl-saoH₂ (0.15 g, 0.60 mmol) and NaOH (0.05 g, 0.75 mmol) were stirred in MeCN (10 cm³) for 2 hours. This slurry solution was then transferred to a 23 cm³ capacity Teflon vial which was then placed in an spring loaded stainless steel digestion vessel and then placed in an oven. The oven temperature was raised to 100 °C over a period of 5 minutes and held at this temperature for 24 hours. The temperature was then lowered gradually down to room temperature over a period of 24 hours. Dark green rhombic shaped crystals of 4 were collected and air dried to give a yield of 12%. Elemental analysis (%) calculated (found) for C₁₆₂H₂₉₈N₉O₃₅Na₃Cr₆: C 57.17 (57.01), H 7.20 (7.45), N 6.06 (6.26). FT-IR (cm⁻¹): 3406(vb), 2957(w), 1589(w), 1549(s), 1484(m), 1460(w), 1420(s), 1378(m), 1361(m), 1296(w), 1277(w), 1255(m), 1229(m), 1201(w), 1169(m), 1020(m), 987(w), 956(w), 898(w), 872(v), 835(m), 812(w), 788(m), 746(w), 715(s), 695(m).

[N₂O₅][NaCr(m)₆(O)₃(O₂C-C₆H₄)₆(Me-sao)₆(MeCN)₂](Na₄Cr₆(O)₂(O₂CC₆H₄)₆(Naphth-sao)₆) (5). Chromium acetate (0.1 g, 0.17 mmol), Me-saoH₂ (0.08 g, 0.53 mmol) and NaOH (0.03 g, 0.75 mmol) were stirred in MeCN (10 cm³) for 1.5 hours. This slurry solution was then transferred to a 23 cm³ capacity Teflon vial which was then placed in an spring loaded stainless steel digestion vessel and then placed inside an oven. The oven temperature was raised to 100 °C over a period of 5 minutes and held at this temperature for 24 hours. The temperature was then lowered gradually down to room temperature over a period of 24 hours. Dark green rhombic shaped crystals of 5 were collected and air dried to give a yield of 15%. Elemental analysis (%) calculated (found) for C₄₃H₇₆O₃₅Na₂Cr₆: C 57.85 (58.14), H 5.45 (5.13), N 3.14 (3.96). FT-IR (cm⁻¹): 3046(vb), 2957(w), 1589(w), 1549(s), 1481(w), 1430(vs), 1301(vs), 1296(w), 1255(m), 1181(m), 1141(w), 1087(w), 1041(m), 1025(m), 951(s), 885(w), 864(w), 822(w), 785(m), 750(m), 735(s), 678(s).

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Notes and references


